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actual yield of a

reaction is 22 g and

the theoretical yield is

25 g. Calculate the

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6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the following equation:

$$N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$$

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ANSWER Answer the following questions in the space provided. 1.

Given the following equation: $C_3H_4(g) + x$.

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$O_2(g) + 3CO_2(g) + 2H_2O(g)$ a. What is the value of the coefficient . x. in this equation? b. What is the molar mass of C_3H_4 ? c. How many moles are in an 8.0 g sample of C_3H_4 ? 2. a. What is meant by . ideal conditions

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visualize chemical rxns

from a math, micro and

macro point of view

...use stoichiometry to

convert moles &/or

grams of one reactant

&/or product into moles

&/or grams of different

reactants &/or

products

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Theoretically, how many moles of NH_3 will be produced?

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25 g. Calculate the percentage yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the ...

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